

Edexcel Chemistry IGCSE 3.16 - Rates of Reaction

Investigate the effect of different solids on the catalytic decomposition of hydrogen peroxide solution

Flashcards

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What is a catalyst?







What is a catalyst?

A substance which increases the rate of a reaction without being chemically changed at the end







How does a catalyst increase the rate of a reaction?







How does a catalyst increase the rate of a reaction?

Provides an alternative pathway which has a lower activation energy.

More collisions will exceed activation energy so more frequent successful collisions.







State 2 things that should be measured when investigating the rate of hydrogen peroxide decomposition







State 2 things that should be measured when investigating the rate of hydrogen peroxide decomposition

Volume of gas produced

Time







What apparatus is required to measure the rate of hydrogen peroxide decomposition with different catalysts?







What apparatus is required to measure the rate of hydrogen peroxide decomposition with different catalysts?

- Conical flask
- Water trough
- Measuring cylinder
- Capillary tube
- Bung





Describe an experiment to measure the rate of hydrogen peroxide decomposition using different catalysts







Describe an experiment to measure the rate of hydrogen peroxide decomposition using different catalysts

- Add H₂O₂ to a conical flask and connect this to an upturned measuring cylinder filled with water using a capillary tube
- Add catalyst to the conical flask and close the bung quickly
- Measure the volume of gas produced over a set time
- Repeat with different catalysts







When measuring the rate of hydrogen peroxide decomposition, why is it important that the bung is connected as soon as the catalyst is added?

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When measuring the rate of hydrogen peroxide decomposition, why is it important that the bung is connected as soon as the catalyst is added?

To ensure minimal gas is lost







Write an equation for the decomposition of hydrogen peroxide







Write an equation for the decomposition of hydrogen peroxide

 $2H_2O_2 \rightarrow O_2 + 2H_2O_2$ ()r $H_2O_2 \rightarrow \frac{1}{2}O_2 + H_2O_2$

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What catalyst is used to increase the rate of hydrogen peroxide decomposition?







What catalyst is used to increase the rate of hydrogen peroxide decomposition?





